**pH and pOH Calculations**

1) Determine the pH of a 0.0034 M HNO3 solution.

2) Determine the pOH of a 0.0034 M HNO3 solution.

3) Determine the pH of a 4.3 x 10-4 M NaOH solution.

4) If a solution is created by adding water to 2.3 x 10-4 moles of NaOH and

4.5 x 10-6 moles of HBr until the final volume is 1 L, what is the pH of this solution?

5) Determine the pH of a 4.5 x 10-11 M NaOH solution.

6) Why would we say that a solution with a H+ concentration of 1.00 x 10-7 M

is said to be neutral. If it contains acid, shouldn’t it be acidic?